

EAMCET CODE: INDI







Sri Indu Institute of Engineering and Technology (Autonomous)

(Formerly RVR Institute of Engineering & Technology)

An Autonomous Institution Under UGC

NAAC Accredited. Recognized Under 2(f) of UGC Act 1956

Approved by AICTE, New Delhi, & Affiliated to JNTUH, Hyderabad.

JNTUH CODE: X3

COURSE FILE

ON

ENGINEERING CHEMISTRY LAB

Course Code - CH206BS

I B. Tech Semester-II A.Y. 2022-2023

Prepared by K.MOUNIKA Asst. Professor

Head of the Department
Department of H&S
SRI INDU INSTITUTE OF ENGG & TECH

heriouda(M) Ibrahimpatham (M) R.R. Dist-501 516

Sri Indu Institute of Engineering & Tech. Sheriguda(Vill), Ibrahimpatnam R.R. Dist. Telangana-501 510.











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Name of the Physical	
laboratory:	ENGINEERING CHEMISTRY LAB
Course code	CH206BS
Room No	D-103 & B-104
Name of the lab incharge	K.MOUNIKA
Name of the faculty incharge	K.MOUNIKA

Index of Lab File

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Main Road, Sheriguda, Ibrahimpatnam, R.R. Dist. 501 510, Telangana. Campus Ph: 9640590999, 9347187999.

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INSTITUTE VISION & MISSION

Vision:

To become a premier institute of academic excellence by providing the world class education that transforms individuals into high intellectuals, by evolving them as empathetic and responsible citizens through continuous improvement.

Mission:

- > **IM1:** To offer outcome-based education and enhancement of technical and practical skills.
- > **IM2:** To Continuous assess of teaching-learning process through institute-industry collaboration.
- ➤ **IM3:** To be a centre of excellence for innovative and emerging fields in technology development with state-of-art facilities to faculty and students' fraternity.

> **IM4:** To Create an enterprising environment to ensure culture, ethics and social responsibility among the stakeholders.

Head of the Department
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PROGRAMME OUTCOMES

PO1: Engineering Knowledge: Apply the knowledge of mathematics, science, engineering fundamentals, and an engineering specialization to the solution of complex engineering problems.

PO2: Problem Analysis: Identify, formulate, research literature, and analyze complex engineering problems reaching substantiated conclusions using first principles of mathematics, natural sciences, and engineering sciences.

PO3: Design / Development of Solutions: Design solutions for complex engineering problems and design system components or processes that meet the specified needs with appropriate consideration for the public health and safety, and the cultural, societal, and environmental considerations.

PO4: Conduct Investigations of Complex Problems: Use research-based knowledge and research methods including design of experiments, analysis and interpretation of data, and synthesis of the information to provide valid conclusions.

PO5: Modern Tool Usage: Create, select, and apply appropriate techniques, resources, and modern engineering and IT tools including prediction and modelling to complex engineering activities with an understanding of the limitations.

PO6: The Engineer & Society: Apply reasoning informed by the contextual knowledge to assess societal, health, safety, legal and cultural issues and the consequent responsibilities relevant to the professional engineering practice.

PO7: Environment & Sustainability: Understand the impact of the professional engineering solutions in societal and environmental contexts, and demonstrate the knowledge of, and need for sustainable development.

PO8: Ethics: Apply ethical principles and commit to professional ethics and responsibilities and norms of the engineering practice.

PO9: Individual & Team Work: Function effectively as an individual, and as a member or leader in diverse teams, and in multidisciplinary settings.

PO10: Communication: Communicate effectively on complex engineering activities with the engineering community and with society at large, such as, being able to comprehend and write effective reports and design documentation, make effective presentations, give and receive clear instructions.

PO11: Project Management & Finance: Demonstrate knowledge and understanding of the engineering and management principles and apply these to one's own work, as a member and leader in a team, to manage projects and in multidisciplinary environments.

PO12: Life-Long Learning: Recognize the need for, and have the preparation and ability to engage in independent and life-long learning in the broadest context of technological change.

Head of the Department
Department of H&S

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SRI INDU INSTITUTE OF ENGINEERING AND TECHNOLOGY

B.Tech. in COMPUTER SCIENCE AND ENGINEERING(AI&ML) COURSE STRUCTURE

I YEAR SYLLABUS (BR22 Regulations)

Applicable from Academic Year: 2022-23 Batch

I Year I Semester

S. No.	Course Code	Course Title	L	T	P	Credits
1.	MA101BS	Matrices and Calculus	3	1	0	4
2.	AP102BS	Applied Physics	3	1	0	4
3.	CS103ES	Programming for Problem Solving	3	0	0	3
4.	ME102ES	Engineering Workshop	0	1	3	2.5
5.	EN104HS	English for Skill Enhancement	2	0	0	2
6.	CS106ES	Elements of Computer Science &	0	0	2	1
		Engineering				
7.	AP105BS	Applied Physics Laboratory	0	0	3	1.5
8.	CS107ES	Programming for Problem Solving Laboratory	0	0	2	1
9.	EN107HS	English Language and Communication Skills	0	0	2	1
		Laboratory				
10.	*MC101ES	Environmental Science	3	0	0	0
11.		Induction Programme				
		Total	14	3	12	20

I Year II Semester

S. No.	Course Code	Course Title	L	Т	P	Credits
1.	MA201BS	Ordinary Differential Equations and Vector Calculus	3	1	0	4
2.	CH203BS	Engineering Chemistry	3	1	0	4
3.	ME201ES	Computer Aided Engineering Graphics	1	0	4	3
4.	EE201ES	Basic Electrical Engineering	2	0	0	2
5.	EC201ES	Electronic Devices and Circuits	2	0	0	2
6.	CH206BS	Engineering Chemistry Laboratory	0	0	2	1
7.	EE202ES	Basic Electrical Engineering Laboratory	0	0	2	1
8.	CS201ES	Python Programming Laboratory	0	1	2	2
9.	CS203ES	IT Workshop	0	0	2	1
		Total	11	3	12	20



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ENGINEERING CHEMISTRY LABORATORY (Course Code: CH206BS)

B.Tech. I Year II Sem.

L T P C
0 0 2 1

Prerequisites: Engineering Chemistry

Course Objectives: The course consists of experiments related to the principles of chemistry required for engineering student. The student will learn:

- Estimation of hardness of water to check its suitability for drinking purpose.
- Students are able to perform estimations of acids and bases using conductometry, potentiometry and pH metry methods.
- Students will learn to prepare polymers such as Bakelite and nylon-6 in the laboratory.
- Students will learn skills related to the lubricant properties such as saponification value, surface tension and viscosity of oils.

Course Outcomes: The experiments will make the student gain skills on:

- Determination of parameters like hardness of water and rate of corrosion of mild steel in various conditions.
- Able to perform methods such as conductometry, potentiometry and pH metry in order to find out the concentrations or equivalence points of acidsand bases.
- Students are able to prepare polymers like bakelite and nylon-6.
- Estimations saponification value, surface tension and viscosity of lubricant oils.

List of Experiments:

- **I. Volumetric Analysis:** Estimation of Hardness of water by EDTA Complexometry method.
- **II. Conductometry:** Estimation of the concentration of an acid by Conductometry.
- **III. Potentiometry:** Estimation of the amount of Fe⁺² by Potentiomentry.
- IV. pH Metry: Determination of an acid concentration using pH meter.

V. Preparations:

- 1. Preparation of Bakelite.
- 2. Preparation Nylon 6.

VI. Lubricants:

- 1. Estimation of acid value of given lubricant oil.
- 2. Estimation of Viscosity of lubricant oil using Ostwald's Viscometer.

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VII. Corrosion: Determination of rate of corrosion of mild steel in the presence and absence of inhibitor.

VIII. Virtual lab experiments

- 1. Construction of Fuel cell and its working.
- 2. Smart materials for Biomedical applications
- 3. Batteries for electrical vehicles.
- 4. Functioning of solar cell and its applications.

REFERENCE BOOKS:

- 1. Lab manual for Engineering chemistry by B. Ramadevi and P. Aparna, S Chand Publications, New Delhi (2022)
- 2. Vogel's text book of practical organic chemistry 5th edition
- 3. Inorganic Quantitative analysis by A.I. Vogel, ELBS Publications. College Practical Chemistry by V.K. Ahluwalia, Narosa Publications Ltd. New De



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COURSE OUTCOMES

Course Name: Engineering Chemistry Lab (C126)

At the End of the course, student will be able to

- C126.1 Determination of parameters like hardness of water by the complexometric titrations (Understanding L2)
- C126.2 Students can able to perform the methods such as conductometry, pH metry to find out concentration of unknown solutions. (Applying L3)
- C126.3 Students can determine the Potentiometry in order to find out the concentrations of acids and bases. (Applying L3)
- C126.4 Students are able to synthesise Polymers-Bakelite & Nylon-6. (Applying L3)
- C126.5 Students can estimate the saponification value and viscosity of the lubricants. (Analyzing L4)
- C126.6 They can able to demonstrate the rate of corrosion of mild steel in various conditions (Understanding L2)

COs and POs & PSOs Mapping

High -3 Medium-2 Low-1

Course Outcome	PO 1	PO 2	PO 3	PO 4	PO 5	PO 6	PO 7	PO 8	PO 9	PO 10	PO 11	PO 12	PSO 1	PSO 2
C126.1	2	2	-	-	-	-	1	-	_	-	-	-	-	-
C126.2	2	2	-	-	1	-		-	_	-	-	-	-	-
C126.3	2	1	-	1	1	-	1	-	_	-	-	-	-	-
C126.4	2	-	2	-	-	1		-	-	-	-	-	-	-
C126.5	2	1	-	-	-		1	-	-	-	-	-	-	-
C126.6	2	1	-	1	-	1	1	-	-	-	-	-	-	-
AVE	2.00	1.4	2.0	1.0	1.0	1.0	1.0	-	-	-	-	-	-	-



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MAPPING OF EXPERIMENT OUTCOMES WITH CO/PO'S/PSO

EXPERIMENT OBJECTIVES	EXPERIMENT OUTCOMES	СО	PO'S
1.To estimate the total hardness of water byEDTA method.	The students will be able to analyze the nature of salts causing hardness and to solve the engineering problems arising during steam production in boilers.	C126.1	PO1, PO2 ,PO7
2.To determine the strength ofthe strong acid bytitration with strong base conductometrically.	The student shall be able to analyze the variation of conductance values of given acid with addition of strong base using conductivity meter.	C126.2	PO1, PO2 PO5
3.To estimate the Fe ⁺² by potentiometry using KMnO ₄ .	The student shall be able to Analyze the variation of EMF values of given acid with addition of KMnO ₄ using potentiometer.	C126.3	PO1, PO2, PO4, PO5, PO7
4.To estimate the amount of HCl present in the given volume of test solution by P ^H metry.	The student shall be able to find out the concentrations of acids and bases.	C126.2	PO1, PO2 PO5
5.To prepare Bakelite polymer using Phenoland Formaldehyde.	The student shall be able to prepare the polymer of Bakelite	C126.4	PO1, PO3 PO6
6.To prepare Nylon-6,6 polymer using adipoyl chloride and hexamethylenediammine.	The student shall be able to prepare the polymer of Nylon-6,6.	C126.4	PO1, PO3 PO6
7.To determine the acid value of Coconut oil.	The student shall be able to identity and calculate the acid values of coconut oil.	C126.5	PO1, PO2 PO7



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8. To determine the viscosity of the given coconut oil and ethanol.	The students shall be able to Determine and calculate the Viscosity coefficient values of coconut oil and ethanol.	C126.5	PO1, PO2 PO7
9.To determine the rate of corrosion of mild steel in acidic medium in the absence and presence of an inhibitor and calculate the efficiency of the Inhibitor.	The student can able to understand the efficiency and function of inhibitor in the Process of corrosion.	C126.6	PO1, PO2 PO4, PO6 PO7
ADDITIONAL EXPERIMENT	r'S		
10.To determine the strength of the weak acid by titration with strong base conductometrically.	The student shall be able to analyze the variation of conductance values of given weak acid with addition of strong base using conductivity meter.	C126.2	PO1, PO2 PO5
11.To determine the surface tension of a given liquid at room temperature using stalagmometer by drop number method.	The student shall be able to determine and calculate the surface tension values of reference liquid and given liquid	C126.5	PO1, PO2 PO7



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Class:AI&ML-A Semester: II W.E.F-03-04-2023 LH:-D-105

	I 9:40- 10:30	II 10:30 - 11:20	III 11:20- 12:10	12:10- 12.45	IV 12.45- 1.35	V 1.35- 2.25	VI 2.25- 3.15	VII 3.15-4.00
MON	F	EC BEE LA	В	L	EC	EDC	BEE	PYTHON(T)
TUE	EDC	ODE	EC	U		PYTHON	LAB	ODE(T)/EC(T)
WED	CA	EG PRACTI	CE	C H	BEE	ODE	EDC	EDC(T)/BEE(T)
THU	BEE	ODE	BEE	п	I	TWS LAB	1	EC(T)/ODE(T)
FRI		EC/BEE LAI	3		ODE EC EDC		LIBRARY	
SAT	BEE	ODE	EC		CAE	G PRACT	TCE	BEE(T)/EDC(T)

Course Code	Course Name	Name of the Faculty	Course Code	Course Name	Name of the Faculty
MA201BS	ODE-Ordinary Differential Equations & Vector Calculus	V.SRINIVAS	CH206BS	EC LAB Engineering Chemistry Laboratory	Dr.D.PREMALATH A/ K.MOUNIKA
CH203BS	EC-Engineering Chemistry	Dr.D.PREMALATHA	EE202ES	BEE LAB-Basic Electrical Engineering Laboratory	K.RAJASHEKAR/S. NISCHALA
ME201ES	CAEG-Computer Aided Engineering Graphics	M.YADHAGIRI	CS201ES	PYTHON Programming Laboratory	M.TEJASWI/P.BAL U
EE201ES	BEE-Basic Electrical Engineering	K.RAJASHEKAR	CS203ES	ITWS-IT Workshop	N.KEERTHI CHANDANA/B.SW ATHI
EC201ES	EDC-Electronic Devices & Circuits	P.ARUNA KUMARI			

Class In-Charge

Time Papie Coordinator

Head of The Department

Sri Indu Institute of Engg. & Tech Main Road, Sheriguda(V) Ibrahimpatnam(M), R.P. Telangana-50



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BR22

Lab External Question paper

Year & Semester: I-II Branch: AI&ML-A

Subject Name: Engineering Chemistry Lab Faculty Name: K.MOUNIKA

EXTERNAL EXAM OUESTION PAPER

- 1. Estimate the total hardness of water bycomplexometric method using EDTA. [CREATING L6]
- 2. Estimate of an HCL by conductometric titration. [EVALUATING L5]
- 3. Estimate of Fe+2 by Potentiometry using by kmno4. [EVALUATING L5]
- 4. Determine the acid concentration by using PH meter. [EVALUATING L5]
- 5. Estimate of an acetic acid byconductometric titration. [EVALUATING L5]
- 6. Determine the viscosity of a given liquid by using Ostwald's viscometer. [EVALUATING L5]
- 7. Write about preparation of Bakelite. [REMEMBARING L1]
- 8. Write about preparation of Nylon 6,6. [REMEMBARING L1]
- 9. Determine the acid value of coconut oil. [REMEMBARING L1]
- 10. Determine the surface tension of a given liquid by using stalagmometer. [REMEMBARING L1]

STOLENGTH TO STOLE

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EC Lab External Time Table Examination Branch

A.Y.: 2022-23 SEM-II

DATE	Day	Branch	Session	HT.No	Total No of Students
21-8-2023	MONDAY	ECE & CIVIL	FN	22X31A0401 TO 22X31A0464 22X31A0101 TO 22X31A0103	67
22-8-2023	TUESDAY	IOT	FN	22X31A6901 TO 22X31A6963	63
23-8-2023	WEDNESDAY	AI&ML-B	FN	22X31A6651 TO 22X31A6697	47
24-8-2023	THURSDAY	AI&ML-A	FN	22X31A6601 TO 22X31A6650	50
25-8-2023	FRIDAY	AI&DS	FN	22X31A7201 TO 22X31A7264	64

FN: 9:40am to 12:25pm

Head of the Department Department of H&S

SRI INDU INSTITUTE OF ENGG & TECH heriouda(M) Ibrahimpatham (M) R.R. Dist-501 516 PRINCIPAL

Sri Indu Institute of Engineering & Tech Sheriguda(VIII), Ibrahimpatnam R.R. Dist. Telangana-501 510.



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EC Lab External Time Table Examination Branch

A.Y:2022-2023 SEM-II

Date	Day	Branch	Session	HT.NO	Total	Remarks	
					No. of Stude nts	Internal examiner	External examiner
21-08- 2023	MON DAY	ECE & CIVIL	FN	22X31A0401 TO 22X31A0464 22X31A0101 TO 22X31A0103	67	O.SUBHASHINI	D.Swathi 7032248997 Asst.prof BIIET
22-08- 2023	TUE SDA Y	IOT	FN	22X31A6901 TO 22X31A6963	63	V.MOUNIKA	Dr.Nagaveni 9959073712 Assoc.prof BIIET
23-08- 2023	WED NES DAY	AI&ML-B	FN	22X31A6651 TO 22X31A6697	47	O.SUBHASHINI	Dr.Rinki kumar 7488730602 Asst.prof BIIET
24-08- 2023	THU RSD AY	AI&ML- A	FN	22X31A6601 TO 22X31A6650	50	V.MOUNIKA	Dr.Litun swain 9489576721 Asst.prof BIIET
25-08- 2023	FRID AY	AI&DS	FN	22X31A7201 TO 22X31A7264	64	K.MOUNIKA	Dr.Shahroora sameen 9149454924 Asst.prof BIIET

FN: 9:40AM to 12:25PM

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LAB OCCUPANCY CHART

ENGINEERING CHEMISTRY LAB

Class: I B.Tech Semister-II LH:B-104

	I 9:40-10:30	II 10:30 -11:20	III 11:20-12:10	12:10- 12.45	IV 12.45-1.35	V 1.35-2.25	VI 2.25-3.15	VII 3.15- 4.00
MON	I BT	TECH II SEM AI&M	L-A	_				
TUE	I B	TECH II SEM AI&I	OS	L II	II			
WED	I	BTECH IISEM EC	Е	N N	I	BTECH II SEM	И ІОТ	
THU		C						
FRI		H						
SAT								

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LAB OCCUPANCY CHART

ENGINEERING CHEMISTRY LAB

Class: I B.Tech Semister-II LH: D-103

	I 9:40-10:30	II 10:30 -11:20	III 11:20-12:10	12:10- 12.45	IV 12.45-1.35	V 1.35-2.25	VI 2.25-3.15	VII 3.15- 4.00
MON	MA	INTAINANCE						
TUE								
WED								
THU					ΙB	TECH II SEM	AI&DS	
FRI]	I BTECH II SEM AI	&ML-A]	BTECH IISEN	И ЕСЕ	
SAT	I	BTECH II SEM IO	Γ		I]	BTECH II SEM	AI&ML-B	

Head of the Department Department of H&S

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ENGINEERING CHEMISTRY LAB

Do's

- 1. Attend all the practical classes with
- a) Observation note book b) Chemistry practical manual c) A neat hand kerchief
- 2. Follow the instructions of your Lecturer carefully.
- 3. Read the experiment perfectly before starting.
- 4. Take the required apparatus and clean them.
- 5. The observations should be in noted in the note book immediately.
- 6. Clean the apparatus immediately after the experiment and return to the concerned lab incharge.
- 7. Do the calculation and get the signature of Lecturer on the observation note book.
- 8. Always throw the pieces of papers, broken glass pieces etc., in a waste basket only.
- 9. The observations and calculations should be recorded neatly in the record book and submit the same of the lecturer.

Don'ts

- 1. Don't handle the apparatus roughly; it leads in damage (or)breakage.
- 2. Don't performthe experiment with incomplete knowledge, it may lead you in confusion.
- 3. Don't use excess amount of chemicals (or) reagents.
- 4. Don't consult your fellow student, if you have doubt in the experiment, consult Lecturer only.
- 5. Don't throw any solid matter in the sink. Don't waste the chemical or reagent.
- 6. Don't taste any chemical and inhale poisonous gases.
- 7. Don't waste the water. If it is not required, the tap must be closed.
- 8. Don't leave the laboratory unless your work bench is cleanand all theapparatus is returned to the attender.



SRI INDU INSTITUTE OF ENGINEERING & TECHNOLOGY Sheriguda (V), Ibrahimpatnam (M), R.R.Dist-501 510

DEPARTMENT OF HUMANITIES AND SCIENCES

NAME: Engin	eering chemidey	B FLOOR PLAN ROO Flo	OM NO: B-104
windowskill is all	Berches 1.	Benches 8	Habe 41
Door 12'.6" [win LOOF] [40CTable-1]	3 	9 10 11 12	Copjabe
15 A 1900	plack bin is	41-1" 22 flat form	1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1 1
10' Tat Jorn	19 EINK	SINE PERM.	25.4" 51.4"
Area(In. Sqm.)= R	DX 14 = 84 140.	29'.8"	13.6° - 3
Area (In. Sft.)= 30	x 2 = 1560,	H	ead of the Department

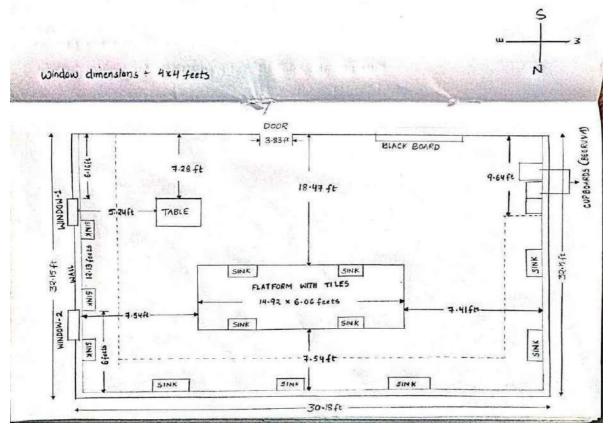


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PHYSICAL LAB FLOOR-1 PLAN

ROOM NO - D103 CHEMISTRY LAB



Lab area in sq.m=9.198*9.80=90.14 Lab area in sq.ft =30.18*32.15=970.287

LAB in Charge

Head of the Department
Department of H&S
SRIINDU INSTITUTE OF ENGG & TECH
Periouda^(N) Ibrahimpamam (N) R.R. Dist-501516



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(Approved by AICTE, New Delhi and Affiliated to JNTUH, Hyderabad)
Khalsa Ibrahimpatnam, Sheriguda (V), Ibrahimpatnam (M), Ranga Reddy Dist., Telangana – 501 510

Lab manual link

https://drive.google.com/file/d/1QjP0tK0jpATEqLV1Wcww6fUdJzJ8mwi9/view?usp=sharing

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anch														
anch		Cor	urse Outcom	e Attainment (Int	ernal Exai	mination	-1)							
anch	f the faculty:	K.MOUNIKA		Academic Year:	_									
	& Section:	AIML-A		Examination:										-
uo ee	ourse Name:	ENGINEERINGC	UEMICTDV	Year/semester		IL I								-
	ourse rvanie.	ENGINEERINGC	HEMISTK I	T car/scinester	1/11									-
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	22X31A6602	10	9	10	<u> </u>									
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	22X31A6614	10	9	10	1	-								-
	22X31A6615	6	8	10										-
-	22X31A6616	10	7	10	1									-
	22X31A6617	10	8	10	1									-
	22X31A6618	9	8	10	1									-
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	22X31A6621	8	5	10	1									
	22X31A6622	10	9	10	i									
23	22X31A6623	10	7	10										
24	22X31A6624	9	7	10										
25	22X31A6625	9	8	10										
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28	22X31A6628	10	10	10										
29	22X31A6629	9	4	10										
	22X31A6630	10	7	10										
-	22X31A6631	10	10	10										
	22X31A6632	10	4	10										
	22X31A6633	8	4	10	ļ									
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	22X31A6635	10	4	10	ļ									<u> </u>
	22X31A6636	9	6	10	 									-
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	22X31A6638	9	7	10	1	-								
	22X31A6639	10	10	10	1	-					-			-
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	22X31A6642 22X31A6643	9	7	10	 	-					-			-
	22X31A6644 22X31A6644	10 A	4 A	10 A	1	+					-			-
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50	75V2TW0020	10	4	10	 	-								-

Target set by the faculty / HoD	6.00	6.00	6.00			
Number of students performed above the target	44	32	47			
Number of students	50	50	50			
Percentage of students cored more than target	88%	64%	94%			
CO Mapping with Exam Qu	uestions:					
CO - 1	Y	V	V			
CO - 2	Y	Y	Y Y			
	+	Y				
CO - 3 CO - 4	Y	Y	Y			
CO - 5	+					
CO - 6						
CO Attainment based on Ex	xam Questions:					
CO - 1	88%	88%	94%			
CO - 2	88%	88%	94%			
CO - 3	88%	88%	94%			
CO - 4	3070	5570	2470			
CO - 5	1					
CO - 6						
СО	Intrnal practica	E+E+R	OveralI	Level	Attainme	nt Level
CO-1	88%	94%	91%	3	1	40%
CO-2	88%	94%	91%	3	2	50%
CO-3	88%	94%	91%	3	3	60%
CO-4						
CO-5	†					
CO-6	†					
	(Internal 1 Ex		<u> </u>			

SRI	INDU INSTI	TUTE OF I	ENGINEER	ING AND TE	CHNOLO									
A.C.	THE ROLL OF THE PARTY OF THE PA	Department of H	Humanities and	Sciences										
16	The state of the s													
1	The state of the s		Course O	utcome Attainme	nt (Internal E	xamination-2)								
Name o	of the faculty:	K,MOUNIKA		Academic Year:	2022-23									
Branch	& Section:	AIML-A		Examination:	INTERNAL	II								
Lab Co	ourse Name:	ENGINEERING (CHEMISTRY	Year/semester	I/II									
S.No	HT No.	R+O+A	V+V	E+E+R	ppt			R+O+.	A: RECORI	+OBSERVA	ATION+ATI	TANDANCE		
	arks ==>	10	10	10	10									
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3	22X31A6602 22X31A6603	10	10 8	10 10	10 10			E.E.I	-EVDEDIME	NT WRITE	IID EVECU	TION DESI	пт	
4	22X31A6604	9	8	9	10			ETETE	LEATERINI	WI WKIIE	UI TEXECU	IONTRES		
5	22X31A6605	9	7	10	10									
6	22X31A6606	9	6	10	10									
7	22X31A6607	10	10	10	10									
8	22X31A6608	10	9 7	10	10									
9	22X31A6609 22X31A6610	10 8	8	10	10 10									
11	22X31A6611	A	A	A	A									
12	22X31A6612	10	9	10	10									
13	22X31A6613	9	8	10	10									
14	22X31A6614	10	9	10	10									
15 16	22X31A6615 22X31A6616	8	7 8	10 10	10 10									
17	22X31A6617	10	10	10	10									
18	22X31A6618	10	10	10	10									
19	22X31A6619	9	7	9	10									
20	22X31A6620	9	7	9	10									
21	22X31A6621	8	6	10	10									
22	22X31A6622 22X31A6623	10	10	10 10	10 10									
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25	22X31A6625	9	8	10	10									
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31	22X31A6631	9	8	10	10									
32	22X31A6632	7	6	7	10									
33	22X31A6633	7	6	7	10									
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40	22X31A6640	A	A	A	10									
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46	22X31A6646	10	6	8	10									
47	22X31A6647	10	6	10	10									
48	22X31A6648	10	7	10 10	10									
49 50	22X31A6649 22X31A6650	10 9	9	10	10									
50	22/01/00/00	,	,	10	10									
		Į.						-						

Target set by the faculty /	6.00	6.00	6.00	6.00				
Number of students performed above the target	48	48 42		49				
Number of students attempted	50	50	50	50				
Percentage of students scored more than target	96%	84%	94%	98%				
CO Mapping with Exam Qu	estions:							
CO - 1								
CO - 2								
CO - 3								
CO - 4	Y	Y	Y	Y				
CO - 5	Y	Y	Y	Y				
CO - 6	Y	Y	Y	Y				
CO Attainment based on Ex	am Questions:							
CO - 1								
CO - 2								
CO - 3								
CO - 4	96%	84%	94%	94%				
CO - 5	96%	84%	94%	94%				1
CO - 6	96%	84%	94%	94%				
СО	Intrnal practica	E+E+R	ppt	OveralI	Level		Attainmer	nt Level
CO-1	F 33350	_		7			1	40%
CO-2							2	50%
CO-3							3	60%
CO-4	90%	94%	94%	93%		3		- 33,0
			94%	93%		3		
ICO-5	90%	94%						
CO-5 CO-6	90%	94%	94%	93%		3		



Department of Humanities and Sciences

Evinis !	1	Department of Hu				
West of the second		Course Outcome Attainment	(Univ	ersity Exa	<u>minations)</u>	
Name of	of the faculty:	K.MOUNKA		Academic	Year:	2022-23
Branch	& Section:	AIML-A		Year / Sen	nester:	I/II
Lab C	ourse Name:	ENGINEERING CHEMISTRY				
S.No	Roll Number	Marks Secured		S.No	Roll Number	Marks Secured
1	22X31A6601	58		35	22X31A6635	50
2	22X31A6602	60		36	22X31A6636	54
3	22X31A6603	50		37	22X31A6637	50
4	22X31A6604	51		38	22X31A6638	50
5	22X31A6605	55		39	22X31A6639	59
6	22X31A6606	50		40	22X31A6640	47
7	22X31A6607	60		41	22X31A6641	54
8	22X31A6608	58		42	22X31A6642	55
9	22X31A6609	59		43	22X31A6643	35
10	22X31A6610	35		44	22X31A6644	58
11	22X31A6612	55		45	22X31A6645	58
12	22X31A6613	55		46	22X31A6646	56
13	22X31A6614	60		47	22X31A6647	59
14	22X31A6615	52		48	22X31A6648	59
15	22X31A6616	60		49	22X31A6649	55
16	22X31A6617	54		50	22X31A6650	58
17	22X31A6618	57				
18	22X31A6619	35				
19	22X31A6620	53				
20	22X31A6621	35				
21	22X31A6622	60				
22	22X31A6623	58				
23	22X31A6624	53				
24	22X31A6625	56				
25	22X31A6626	60				
26	22X31A6627	60				
27	22X31A6628	57				
28	22X31A6629	52				
29	22X31A6630	60				
30	22X31A6631	60				
31	22X31A6632	52				
32	22X31A6633	45				
33	22X31A6634	55				
Class A	verage mark		54		Attainment Level	% students
		med above the target	32		1	40%
	of successful stude		49		2	50%
Percenta	age of students scor	red more than target	65%		3	60%
	nment level	-	3			
110001						

SRI INDU I	NSTIT	UTE OF EN	GINEE	RING AND T	TECHNOLOGY
The state of the s	Departme	ent of Humanities	and Scier	ices	
Sylva		Course Ou	tcome At	<u>tainment</u>	
West marather					
Name of the faculty	K.MOUN	IIKA		Academic Year:	2022-23
Branch & Section:	AIML-A			Year / Semester:	I/II
Lab Course Name:	ENGINEE	RINGCHEMISTRY			
Course Outcomes	1st Internal Exam	2nd Internal Exam	Internal Exam	University Exam	Attainment Level
CO1	3.00		3.00	3.00	3.00
CO2	3.00		3.00	3.00	3.00
CO3	3.00		3.00	3.00	3.00
CO4		3.00	3.00	3.00	3.00
CO5		3.00	3.00	3.00	3.00
CO6		3.00	3.00	Academic Year: 2022-23 Year / Semester: I/II University Exam Attainment I 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00 3.00	3.00
Inter	nal & Univ	ersity Attainment:	3.00	3.00	
		Weightage	70%	30%	
CO Attainment for the	course (In	ternal, University)	2.10	0.90	
CO Attainment for	the course ((Direct Method)		3.00	
Overall co	urse a	attainmer	nt leve	el	3.00

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SVI.					partn							es										
3,11,3	7			rogra	am Ou	tcome	Attan	ıment	(fron	1 Cour	se)											
ame of	Facu	ltv:	K.MC	UNIK	A			Acad	emic `	Year:	2022	-23										
ranch 8			AIML					_		ester:	_											
ourse N	lame		ENGI	NEERII	NG CHE	MISTE	RY															
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